1. The CORRECT order of pK_a for the compounds I to IV in water at 298 K is

$HCo(CO)_4$	$HCo(CO)_3(CO)_3$	$HCo(CO)_3(P(Oh)_3$	$HCo(CO)_2(PPh_3)_2$
I	II	III	IV

(A) I > II > II > IV
(B) IV > III > II >I
(C) IV > II > III > I
(D) I > III > II > IV

2. For Na^+, M^{2+} > Na^+ >F^- , the CORRECT order of ionic radu is

(A) Al^{3+} > Mg^{2+} >Na^+ > F^-
(B) Al^{3+} > Na^+ >Mg^{2+} > F-
(C) F^- > Na^+ >Mg^{2+} > Al^{3+}
(D) Na^+ >F^- >Mg^{2+} > Al^{3+}

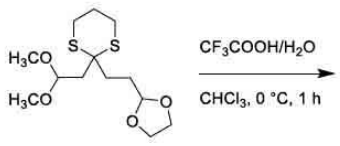
3. Spin-only magnetic moments (in BM) of $[NiCl_2(PPh_3)_2]$ and $[Mn(NCS)_6]^{4-}$, respectively, are

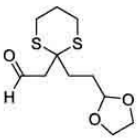
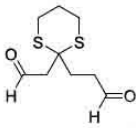
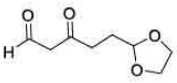
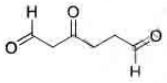
(A) 0.00 and 5.92
(B) 2.83 and 1.89
(C) 0.00 and 1.89
(D) 2.83 and 5.92

4. Two sets of quantum numbers with the same number of radial nodes are

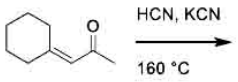
(A) $n = 3; l = 0; ml = 0 \; and \; n = 2; l = 0; ml = 0$
(B) $n = 3; l = 1; ml = 1 \; and \; n = 2; l = 1; ml = 0$
(C) $n = 3; l = 2; ml = 0 \; and \; n = 2; l = 1; ml = 0$
(D) $n = 3; l = 1; ml = -1 \; and \; n = 2; l = 1; ml = 0$

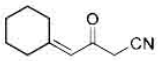
5. The major product formed in the following reaction is

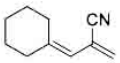


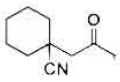
(A) 
(B) 
(C) 
(D) 

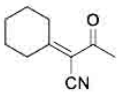
6. The Major product formed in the following reaction is



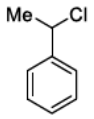
(A) 

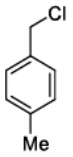
(B) 

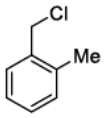
(C) 

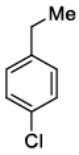
(D) 

7. A compound shows ^{1}H NMR peaks at δ-values (in ppm) 7.31 (2H), 7.21 (2H), 4.5 (2H) and 2.3 (3H). The structure of the compound is

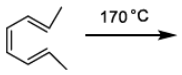
(A) 

(B) 

(C) 

(D) 

8. The major product formed in the following reaction is



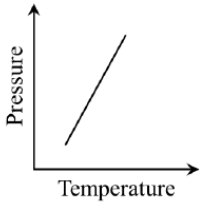
(A) 

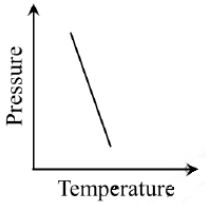
(B) 

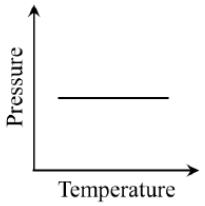
(C) 

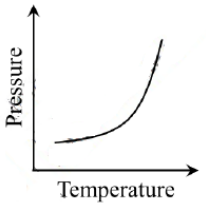
(D) 

9. A pure substance M has lesser density in solid state than in liquid state. The ΔS fusion of M is +25 J K^{-1} mol^{-1}. The CORRECT representative Pressure-Temperature diagram for the fusion of M is

(A) 

(B) 

(C) 

(D) 

10. Among the following , the matrices with non-zero determinant are

$$P=\begin{bmatrix}1&0&0&0\\0&1&0&0\\0&0&1&0\\0&0&0&1\end{bmatrix}\quad Q=\begin{bmatrix}1&0&0&0\\0&2&0&0\\0&0&3&0\\0&0&0&4\end{bmatrix}\quad R=\begin{bmatrix}1&0&0&0\\2&2&0&0\\3&1&3&0\\4&3&1&4\end{bmatrix}\quad S=\begin{bmatrix}1&2&3&1\\2&3&4&2\\3&4&1&3\\4&1&2&4\end{bmatrix}$$

(A) P, Q and R
(B) P, R and S
(C) P, Q and S
(D) Q, R and S

11. Reaction of BCl_3 with NH_4Cl at 140 ^{0}C produces compound P. Further, P reacts with $NaBH_4$ to give a colorless liquid The reaction of Q with H_2O at 100 ^{0}C produces compound R and a diatomic gas S. Among the following, the CORRECT statement is

(A) P is $B_3N_3H_6$
(B) R is $[B(OH)NH]_3$
(C) Q is $[BClNH]_3$
(D) S is Cl_2

12. The complex that does NOT obey the 18-electron rule is (Given: Atomic numbers of Ti, Mn, Ta and Ir are 22, 25, 73 and 77, respectively)

(A) $[(\eta^5\text{-}C_5H_5)Ti(CO)_4]^-$
(B) $[Mn(SnPh_3)_2(CO)_4]^-$
(C) $[(\eta^5\text{-}C_5H_5)Ir(CH_2)(PMe_3)]$
(D) $[TaCl_3(PEt_3)_2(CHCMe_3)]$

13. Hybridization of the central atoms in I_3–, ClF_3 and SF_4, respectively, are

(A) sp^3d, sp^2 and dsp^2
(B) sp, sp^3d and dsp^2
(C) sp^3d, sp^3d and sp^3d
(D) sp, sp^2 and sp^3d

14. Reaction of $[Ni(CN)_4]^{2-}$ with metallic potassium in liquid ammonia at −33 ^{0}C yields complex E. The geometry and magnetic behavior of E, respectively, are

(A) Square planar and diamagnetic
(B) Tetrahedral and diamagnetic
(C) Octahedral and paramagnetic
(D) Square pyramidal and paramagnetic

15. The decreasing order of C=C bond length in the following complexes is

$[Cl_3Pt(CH_2{=}CH_2)]^-$ I $[Cl_3Pt(C(CN)_2{=}C(CH)_2)]^-$ II $[Cl_3Pt(CF_2{=}CH_2)]^-$ III $[Cl_3Pt(CF_2{=}CF_2)]^-$ IV

(A) II > III > IV > I
(B) IV > II > III > I
(C) II > IV > III > I
(D) IV > II > I > III

16. The CORRECT combination for metalloenzymes given in Column I with their catalytic reactions in Column II is

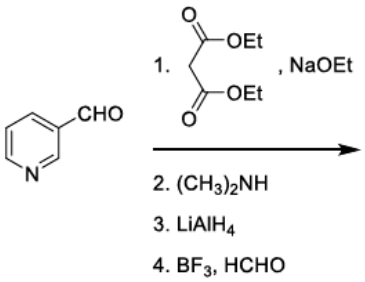
Column I	**Column II**
(i) Cytochrome P-450	(K) $2H_2O_2 \rightarrow 2H_2O+O_2$
(ii) Catalase	(L) $R\text{-}CH_2OH+O_2 \rightarrow R\text{-}CHO+H_2O_2$ (R=alkyl or aryl)
(iii) Galactose oxidase	(M) $O_2+4H^++4e^- \rightarrow 2H_2O$
(iv) Cytochrome c oxidase	(N) $R\text{-}H+O_2+2e^-+2H^+ \rightarrow R\text{-}OH+H_2O$ (R=alkyl or aryl)

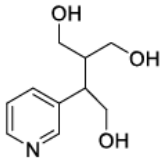
(A) (i)–(M); (ii)–(N); (iii)–(K); (iv)–(L)
(B) (i)–(N); (ii)–(L); (iii)–(K); (iv)–(M)
(C) (i)–(N); (ii)–(K); (iii)–(L); (iv)–(M)
(D) (i)–(M); (ii)–(K); (iii)–(L); (iv)–(N)

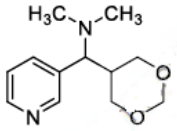
17. According to the crystal field theory, $d-d$ transition observed in $[Ti(H_2O)_6]^{3+}$ is

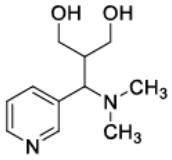
(A) Laporte forbidden and spin forbidden
(B) Laporte allowed and spin forbidden
(C) Laporte allowed and spin allowed
(D) Laporte forbidden and spin allowed

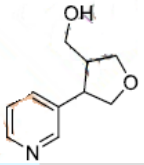
18. The major product formed in the following reaction sequence is



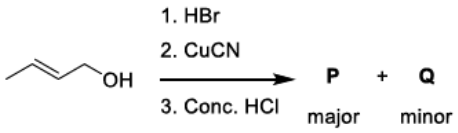
(A) 

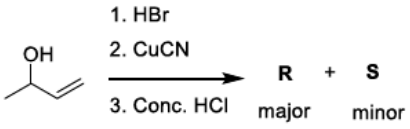
(B) 

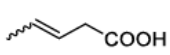
(C) 

(D) 

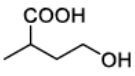
19. The products P,Q,R and S formed in the following reaction are

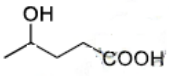


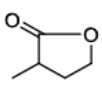


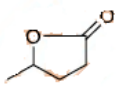
(A) **P = R =** 

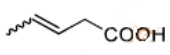
and **Q = S =** 

(B) **P = S =** 

and **Q = R =** 

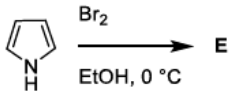
(C) **P = S =** 

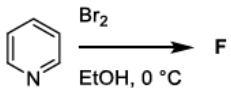
and **Q = R =** 

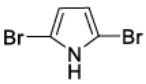
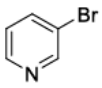
(D) **P = R =** 

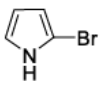
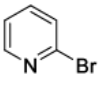
and **Q = S =** 

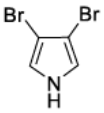
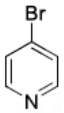
20. The major product E and F formed in the following reactions are

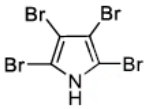
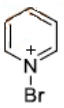




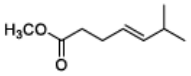
(A) **E** =  and **F** = 

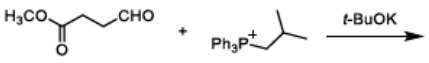
(B) **E** =  and **F** = 

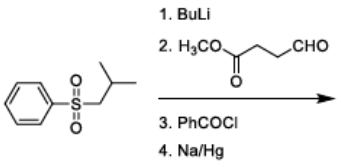
(C) **E** =  and **F** = 

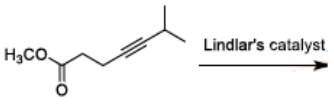
(D) **E** =  and **F** = 

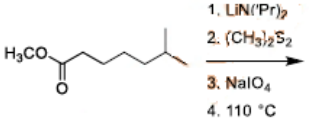
21. The reaction that product the following as a major product is



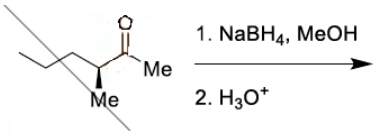
(A) 

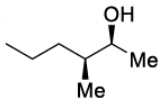
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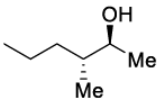
(C) 

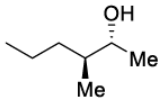
(D) 

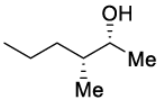
22. The major product formed in the following reaction is



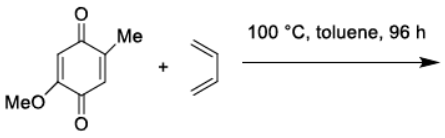
(A) 

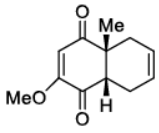
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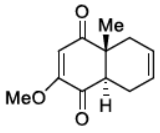
(C) 

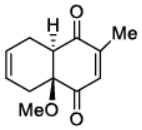
(D) 

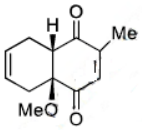
23. The major product formed in the following reaction is



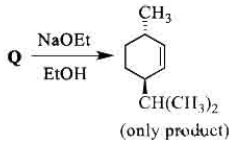
(A) 

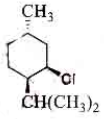
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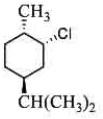
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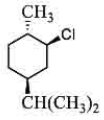
(D) 

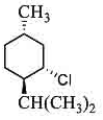
24. In the following reaction, compound Q is



(A) 

(B) 

(C) 

(D) 

25. Monochromatic X-rays having energy $2.8 \times 10^{-15}\ J$ diffracted (first order) from (200) plane of a cubic crystal at an angle 8.5^0. The length of unit cell in Å of the crystal (rounded off to one decimal place) is (Given: Planck's constant, $h = 6.626 \times 10^{-34}\ J\ s$; $c = 3.0 \times 10^{8}\ m\ s^{-1}$)

(A) 2.4
(B) 3.4
(C) 4.8
(D) 9.8

26. For $\alpha > 0$, the value of the ingral $\int_{-\infty}^{+\infty} xe^{-\alpha x^2} dx$ is

(A) $\sqrt{\frac{\pi}{\alpha}}$
(B) ∞
(C) 0
(D) 1

27. The volume correction factor for a non-ideal gas in terms of critical pressure (p_c), critical molar volume (V_c)critical temperature (T_c) and gas constant (R) is

(A) $\frac{RT_c}{8p_c}$
(B) $\frac{27R^2Tc^2}{64p_c}$
(C) $\frac{8P_cV_c}{3T_c}$
(D) $3p_cV_c^2$

28. Half-life ($t_{1/2}$) of a chemical reaction varies with the initial concentration of reactant (A_o) as given below

A_0(mol L^{-1})	5×10^{-2}	4×10^{-2}	3×10^{-2}
$t_{1/2}$ (s)	360	450	600

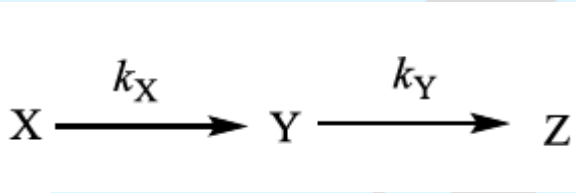
The order of the reaction is

(A) 0
(B) 1
(C) 2
(D) 3

29. The CORRECT statement regarding the molecules BF_3 and CH_4 is

(A) Both BF_3 and CH_4 are microwave active
(B) Both BF_3 and CH_4 are infrared active
(C) CH_4 is microwave active and infrared inactive
(D) BF_3 is microwave active and infrared active

30. For the consecutive reaction,



C_0 is the initial concentration of of X the concentration of X, Y and Z at time t are C_X, C_Y and C_Z, respectively. The expression for the concentration Y at time t is

(A) $\frac{k_XC_o}{k_Y-k_X}\left(e^{-k_yt} - e^{-k_Yt}\right)$
(B) $\frac{k_XC_X}{k_Y-k_X}\left(e^{-k_Xt} - e^{-k_Yt}\right)$
(C) $\frac{k_XC_o}{k_Y-k_X}\left(e^{-k_yt} - e^{-k_Xt}\right)$
(D) $\frac{k_XC_X}{k_Y-k_X}\left(e^{-k_yt} - e^{-k_Xt}\right)$

31. The CORRECT statement(s) about the species is (are)

(A) $CpMo(CO)_3$ and $CpW(CO)_3$ are isoelectronic (where Cp is cyclopentadienyl)
(B) CH_2- and NH_2 are isolobal and isoelectronic
(C) BH and CH are isolobal and isoelectronic
(D) CH_3 and $Mn(CO)_5$ are isolobal

32. The complex(es) that show(s) Jahn-Teller distortion is (are)

(A) $[Co(CN)_5(H_2O)]^{3-}$

(B) $[NiF_6]^{2-}$

(C) $[Mn(CNMe)_6]^{2+}$

(D) $[Co(en)_2F_2]^{+}$

33. The CORRECT statement(s) about sodium nitroprusside is (are)

(A) It is a paramagnetic complex

(B) Nitroprusside ion is formed in the brown ring test for nitrates

(C) It is used for the detection of S^{2-} in aqueous solution

(D) It contains nitrosyl ligand as NO^{+}

34. The pigment responsible for red color in tomato has one functional group. The CORRECT statement(s) about this functional group is (are)

(A) It decolorizes bromine water

(B) It gives hydrazone derivative on reaction with 2,4-dinitrophenylhydrazine

(C) It gets cleaved on reaction with ozone

(D) It gives positive silver mirror test

35. Hantzsch pyridine synthesis involves several steps. Some of those are

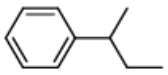
(A) Aldol reaction

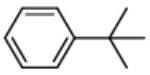
(B) Darzens reaction

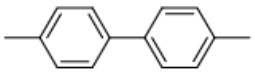
(C) Mannich reaction

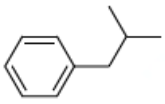
(D) Michael addition

36. The compound(s), which give(s) benzoic acid on oxidation with $KMnO_4$, is (are)

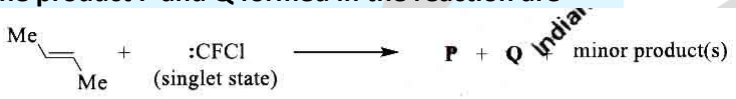
(A) 

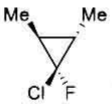
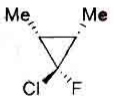
(B) 

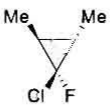
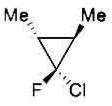
(C) 

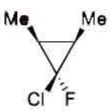
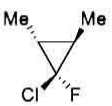
(D) 

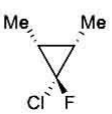
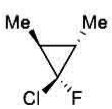
37. The product P and Q formed in the reaction are



(A) P =  Q = 

(B) P =  Q = 

(C) P =  Q = 

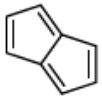
(D) P =  Q = 

38. The functional group(s) in reducing sugar that tests positive with Tollen's reagent is (are)

(A) Aldehyde
(B) Ketone
(C) Hemi-acetal
(D) Acetal

39. Among the following, the anti-aromatic compound(s) is (are)

(A) 

(B) 

(C) 

(D) 

40. The CORRECT Maxwell relation(s) derived from the fundamental equations of thermodynamics is (are)

(A) $\left(\frac{\partial s}{\partial p}\right)_T = -\left(\frac{\partial V}{\partial T}\right)_p$

(B) $\left(\frac{\partial s}{\partial V}\right)_T = \left(\frac{\partial p}{\partial T}\right)_V$

(C) $\left(\frac{\partial T}{\partial V}\right)_s = \left(\frac{\partial p}{\partial s}\right)_V$

(D) $\left(\frac{\partial T}{\partial p}\right)_s = \left(\frac{\partial V}{\partial s}\right)_p$

41. The total number of optically active isomers of dichloridobis (glycinato) cobaltate (III) ion is _____.

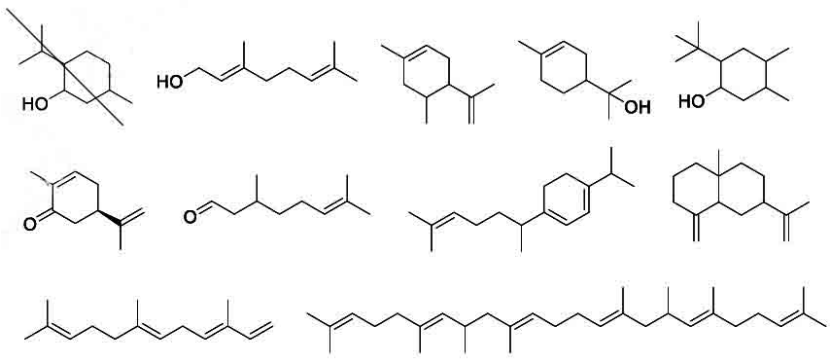
42. The total number of microstates possible for a d^8 electronic configuration is _____.

43. For the following fusion reaction, $4\ {}^1H \rightarrow {}^4He + 2\beta^+ 2v + \gamma$ the Q −value (energy of the reaction) in MeV (rounded off to one decimal place) is _____. (Given: Mass of 1H nucleus is 1.007825 u and mass of ^{4}He nucleus is 4.002604 u)

44. MgO crystallizes as rock salt structure with unit cell length 2.12 Å. From electrostatic model, the calculated lattice energy in kJ mol^{-1} (rounded off to the nearest integer) is _____. (Given: N_A = 6.022 ×1023 mol^{-1}; Madelung constant = 1.748; ε_0 = 8.854 × 10^{-12} J^{-1} C2m^{-1}; charge of an electron = 1.602 × 10^{-19} C)

45. Calcium crystallizes in fcc lattice of unit cell length 5.56 Å and density 1.4848 g cm^{-3} . The percentage of Schottky defects (rounded off to one decimal place) in the crystal is _____. (Given: Atomic mass of Ca is 40 g mol^{-1}; N_A = 6.022× 10^{23} mol^{-1})

46. Among the following, the total number of terpenes(terpenoids) is _____.



47. A buffer solution is prepared by mixing 0.3 M NH_3 and 0.1 M NH_4NO_3. If K_b of NH_3 is 1.6×10^{-5} at 25 ^{0}C, then the pH (rounded off to one decimal place) of the buffer solution at 25 ^{0}C is _____.

48. The dissociation constant of a weak monoprotic acid is 1.6×10^{-5} and its molar conductance at infinite dilution is 360.5×10^{-4} mho m^2 mol^{-1}. For 0.01 M solution of this acid, the specific conductance is $n \times 10^{-2}$ mho m^{-1}. The value of n (rounded off to two decimal places) is _____.

49. Adsorption of a toxic gas on 1.0 g activated charcoal is 0.75 cm^3 both at 2.5 atm, 140 K and at 30.0 atm, 280 K. The isosteric enthalpy for adsorption of the gas in kJ mol^{-1} (rounded off to two decimal places) is _____. (Given: $R = 8.314\ \text{J K}^{-1}\ \text{mol}^{-1}$)

50. If the root mean square speed of hydrogen gas at a particular temperature is 1900 m s^{-1}, then the root mean square speed of nitrogen gas at the same temperature, in m s^{-1} (rounded off to the nearest integer), is _____. (Given: atomic mass of H is 1 g mol^{-1}; atomic mass of N is 14 g mol^{-1})

51. If the crystal field splitting energy of $[Co(NH_3)_4]^{2+}$ is 5900 cm^{-1}, then the magnitude of its crystal field stabilization energy, in kJ mol^{-1} (rounded off to one decimal place),is _____.

52. A salt mixture (1.0 g) contains 25 wt% of $MgSO_4$ and 75 wt% of M_2SO_4. Aqueous solution of this salt mixture on treating with excess $BaCl_2$ solution results in the precipitation of 1.49 g of $BaSO_4$. The atomic mass of M in g mol^{-1} (*rounded off to two decimal places*) is _____.(Given: the atomic masses of Mg, S, O, Ba and Cl are 24.31, 32.06, 16.00, 137.33 and 35.45 g mol^{-1}, respectively)

53. The intensity of a monochromatic visible light is reduced by 90% due to absorption on passing through a 5.0 mM solution of a compound. If the path length is 4 cm, then the molar extinction coefficient of the compound in M^{-1} cm^{-1} is _____.

54. The surface tension (γ) of a solution, prepared by mixing 0.02 mol of an organic acid in 1 L of pure water, is represented as $\gamma^* - \gamma = A\,log\,(1 + Bc)$ γ^* is the surface tension of pure water, A = 0.03 N m^{-1}, B = 50 mol^{-1} L and c is concentration in mol L^{-1}. The excess concentration of the organic acid at the surface of the liquid, determined by Gibbs adsorption equation at 300 K is $n \times 10^{-6}$ mol m^{-2}. The value of n (rounded off to two decimal places) is _____. (Given: R = 8.314 J K^{-1} mol^{-1})

55. The separation of energy levels in the rotational spectrum of CO is 3.8626 cm^{-1}. The bond length (assume it does not change during rotation) of CO in Å (rounded off to two decimal places) is _____. (Given: Planck's constant h = 6.626×10^{-34} J s; $N_A = 6.022 \times 1023$mol^{-1}; atomic mass of C is 12 g mol^{-1}; atomic mass of O is 16 g mol^{-1}; $c = 3 \times 10^8$ m s^{-1})

56. A dilute solution prepared by dissolving a nonvolatile solute in one liter water shows a depression in freezing point of 0.186 K. This solute neither dissociates nor associates in water. The boiling point of the solution in K (rounded off to three decimal places) *is* _____. (Given: For pure water, boiling point = 373.15 K; cryoscopic constant = 1.86 K (mol kg^{-1})$^{-1}$; ebullioscopic constant = 0.51 K (mol kg^{-1})$^{-1}$)

57. The thermodynamic data at 298 K for the decomposition reaction of limestone at equilibrium is given below $CaCO_3(s) \rightleftharpoons CaO(s) + CO_2(g)$

Thermodynamic quantity	$CaCO_3$(s)	CaO(s)	CO_2(g)
μ^0(kJ mol^{-1})	-1128.8	-604.0	-394.4
ΔH_f^o (kJ mol^{-1})	-1206.9	-635.1	-393.5

The partial pressure of CO_2(g) in atm evolved on heating limestone (rounded off to two decimal places) at 1200 K is _____. (Given: R = 8.314 J K^{-1} mol^{-1})

58. The mean ionic activity coefficient of 0.004 molal $CaCl_2$ in water at 298 K (rounded off to three decimal places) is _____. (Given: Debye-Hückel constant for an aqueous solution at 298 K is 0.509 $kg^{1/2}$ $mol^{-1/2}$)

59. For the reaction,

$$Q + R \underset{k_{-1}}{\overset{k_1}{\rightleftharpoons}} X \xrightarrow{k_2} P$$

$k_1 = 2.5 \times 10^5\ L\ mol^{-1} s^{-1}, k^{-1} = 1.0 \times 10^4\ s^{-1}$ and $k_2 = 10\ s^{-1}$. Under steady state approximation, the rate constant for the overall reaction in L mol^{-1} s^{-1} (rounded off to the nearest integer) is _____.

60. For the molecule, $CH_3–CH=CH–CH(OH)–CH=CH–CH=C(CH_3)_2$ the number of all possible stereoisomers is _____.

ANSWER - KEY

Q.No.	Answer	Q.No.	Answer	Q.No.	Answer
1	C	21	B	41	6
2	C	22	C	42	45
3	D	23	A	43	26.7 to 26.8
4	C	24	C	44	-7880 to -6150 or 6150 to 7880
5	A	25	C	45	3.9 to 4.1
6	C	26	C	46	7
7	B	27	A	47	9.7
8	B	28	C	48	1.44 to 1.47
9	B	29	B	49	-5.81 to -5.75
10	A	30	A	50	507 to 510
11	B	31	A, B, D	51	-84.8 to -84.2 or 84.2 to 84.8
12	D	32	A, C	52	38.98 to 39.25
13	C	33	C, D	53	50
14	B	34	A, C	54	2.60 to 2.62
15	C	35	A, D	55	1.12 to 1.14
16	C	36	A, D	56	373.201
17	D	37	B	57	4.20 to 4.35
18	B	38	A, B, C	58	0.772 to 0.775
19	D	39	B, C, D	59	250
20	D	40	A, B, D	60	8